CHEMISTRY BASIC CONCEPTS MCQs

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BASIC CONCEPTS

MCQs

Q.1 Smallest particle of an element which may or may not have independent
existence
(a) a molecule (b) an atom
(c) an ion (d) an electron

Q.2 Swedish chemist J. Berzelius determined the

(a) atomic no. (b) atomic volume

(c) atomic mass (d) atomic density

Q.3 The number of atoms present in a molecule determine its

(a) molecularity (b) basicity

(c) acidity (d) atomicity

Q.4 When an electron is added to a unipositive ion we get

(a) anion (b) cation

(c) neutral atom (d) molecule

Q.5 CO^+ is an example of:

(a) free radical (b) cationic molecular ion

(c) an ionic molecular ion

(d) stable molecule

Q.6 Relative atomic mass is the mass of an atom of an element as compared to

the mass of

(a) oxygen (b) hydrogen

(c) nitrogen (d) carbon

Q.7 Isotopes are the sister atoms of the same element with similar chemical

properties and different

(a) atomic number (b) atomic mass

(c) atomic volume (d) atomic structure

Q.8 The instrument which is used to measure the exact masses of different

isotopes of an element called

(a) I.R. Spectrophotometer (b) U.V. Spectrophotometer

(c) Mass Spectrometer (d) Colourimeter

Q.9 Mass spectrometer separates different positive isotopic ions on the basis

of their

(a) mass value (b) m/e value
(c) e/m value (d) change value

Q.10 Simplest formula that gives us information about the simple ratio of
atoms in a compound is called
(a) structural formula (b) molecular formula
(c) empirical formula (d) molar ratio
Q.11 Percentage of oxygen in H_2O is
(a) 80% (b) 88.8%
(c) 8.8% (d) 9.8%
Q.12 More abundant isotope of an element is one with
(a) even atomic no. (b) odd atomic no.
(c) Even mass no. (d) odd mass no.
Q.13 Large no. of isotopes are known for the elements whose masses are
multiple of
(a) two (b) four
(c) six (d) eight
Q.14 When 0.01 kg of $CaCO_3$ is decomposed the CO_2 produced occupies a
volume at S.T.P.
(a) 2.2414 dm3 (b) 22.414 dm3
(c) 22414 dm3 (d) 224014 dm3
Q.15 The no. of covalent bond in 10gm of NH_3 are
(a) 6.022 x 1023 (b) 1.062 x 1023
(c) 10.62 x 1024 (d) 1.062 x 1024
Q.16 No. of molecules present in 10gm of water are
(a) 3.37 x 1023 (b) 33.7 x 1023
(c) 3.37 x 1024 (d) 3.037 x 1024

Q.17 The no. of covalent bonds present in 10gm of water are

(a) 6.074×10^{23} (b) 6.74×10^{23}

(c) 6.074×10^{24} (d) 6.74×10^{24}

Q.18 The least no. of molecules present in 30 gm of

(a) N_2O (b) NO

(c) NO_2 (d) N_2O_3

Q.19 Which of the following has highest percentage of nitrogen

(a) $(NH_4)_2SO_4$ (b) $NH_4H_2PO_4$

(c) $(NH_4)_2HPO_4$ (d) $(NH_4)_3PO_4$

Q.20 0.1 mole of Na_3PO_4 completely dissociates in water to produce Na^+

(a) 6.02×10^{22} (b) 6.02×10^{23}

(c) 1.806×10^{23} (d) 1.806×10^{22}

Q.21 Efficiency of chemical reaction can be checked by calculating

(a) amount of limiting reactant

(b) amount of the reactant in excess

(c) amount of the product formed

(d) amount of the reactant unused

Q.22 A limiting reactant is one

(a) which is present in least amount

(b) which produces minimum no. of moles of product

(c) which produces maximum no. of moles of product

(d) does not effect the amount of product

Q.23 Stoichiometry is the branch of chemistry which deals with the study of

quantitative relationship among the various

(a) reactants (b) products

(c) Reactants and products (d) all of above

Q.24 500 cm3 of H2 gas at STP contradictions of hydrogen

(a) 6.02 x 1023 (b) 3.01 x 1022

(c) 2.68 x 1022 (d) 1.34 x 1022

Q.25 Largest number of H+ ions are produced by complete ionization of

(a) 0.01 mole of HCl (b) 0.0050 mole of H2SO4

(c) 0.000334 moles of H3PO4

(d) all above

Q.26 The Avogadro's number is

(a) 6.02 x 1024 (b) 6.02 x 10−24

(c) 6.02 x 10−23 (d) 6.02 x 1023

Q.27 The largest number of H+ are produced by complete ionization of

(a) 0.100 2 moles of HCl (b) 0.051 moles of H2SO4

(c) 0.0334 moles of H3PO4 (d) All of the above

Q.28 A sample of pure matter is

(a) element (b) compound

(c) substance (d) mixture

Q.29 nm stands for

(a) Newton meter (b) Nanometer

(c) Newton square meter (d) none of the above

Q.30 One calorie is equal to

(a) 4.184 J (b) 41.84 J

(c) 0.4184 J (d) 0.04184 J

Q.31 The number of moles of CO2 which contains 8.0 gm of oxygen

(a) 0.25 (b) 0.50

(c) 1.0 (d) 1.50

Q.32 27 grams of Al will react completely with how much mass of O2 to

produce Al2O3

(a) 8 gm of oxygen (b) 16 gm of oxygen

(c) 32 gm of oxygen (d) 24 gm of oxygen

Q.33 Mole of SO_2 contains

(a) 6.02×10^{23} atoms of oxygen

(b) 18.1×10^{23} molecules of SO_2

(c) 6.023×10^{23} atom of sulphur

(d) 4 gram of SO_2

Q.34 The largest number of molecules are presenting

(a) 3.6 gram of H_2O (b) 4.8 gram of C_2H_5OH

(c) 2.8 gm of CO (d) 5.4 gms of N_2O_5

Q.35 The mass of one mole of electron is

(a) 1.008 mg (b) 0.184 mg

(c) 1.673 mg (d) 0.55 mg

Q.36 Isotopes differ in

(a) properties which depend on mass

(b) arrangements of electrons in orbital

(c) chemical properties

(d) the extent to which they may be affected in electromagnetic field

Q.37 The volume occupied by 1.4 gm of N_2 at STP is

(a) 224 dm3 (b) 22.4 dm3

(c) 1.12 dm3 (d) 112 cm3

Q.38 Many elements have fractional atomic mass. This is because

(a) the mass atom is itself fractional

(b) atomic masses are average masses of isobars

(c) atomic masses are averages masses of isotopes

(d) atomic masses are average masses of isotopes proportional to relative abundance

Q.39 A limiting reactant is one which

(a) is taken in lesser quantity in grams as compared to other

reactants

(b) is taken in lesser quantity in volume as compared to the other

(c) gives the maximum amount of the product which is required

(d) gives the minimum amount of the product under consideration

Q.40 Isotopes when even atomic masses are a comparatively abundant

(a) demper's spectrograph is superior to that of Aston's

(b) 0.1 mg of H_2O has greater number of molecules then 0.1 mg of

CH_4

(c) the number of H^+ and PO^{-3} ions are not equal but the number of

positive and negative charges

(d) are equal when 100 molecules of H_3PO_4 are thrown in excess of

water

Q.41 A molecule having two atoms is called

(a) monoatomic molecules (b) diatomic molecules

(c) Polyatomic molecules (d) homoatomic molecule

Q.42 An ordinary misoscope is used to measure the object of size

(a) upto 500 nm (b) upto 850 nm

(c) upto 1000 nm (d) upto 1200 nm

Q.43 1 atomic masses unit (amu) is equation

(a) 1.66×10^{-27} kg (b) 1.56×10^{-27} kg

(c) 1.76×10^{-21} kg (d) 1.8×10^{-27} kg

Q.44 Nickel has isotopes

(a) 1 (b) 3

(c) 5 (d) 7

Q.45 Cadmium has isotopes

(a) 3 (b) 5

(c) 7 (d) 9

Q.46 The pressure of vapours in the separating isotopes by mass spectrometry

is kept at

(a) $10-6$ torr (b) $10-4$ torr

(c) $10-3$ torr (d) $10-5$ torr

Q.47 Number of gram atoms in 0.1 gm of Na is

(a) 0.0043 (b) 0.0403

(c) 0.403 (d) None of these

Q.48 Molecule of haemoglobin contains atoms

(a) 15,000 (b) 12,000

(c) 10,000 (d) 8,000

Q.49 Haemoglobin is heavier than a hydrogen atom

(a) 65,000 (b) 68,000

(c) 62,000 (d) 60,000

Answers

Questions 1 2 3 4 5
Answers b C d c b
Questions 6 7 8 9 10
Answers d b c b c
Questions 11 12 13 14 15
Answers b c b a d
Questions 16 17 18 19 20
Answers a b d d c

Questions 21 22 23 24 25
Answers c b d c d
Questions 26 27 28 29 30
Answers d d a b a
Questions 31 32 33 34 35
Answers a d c a d
Questions 36 37 38 39 40
Answers a c d d c
Questions 41 42 43 44
Answers c a a c
Questions 45 46 47 48 49
Answers d a a c b

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