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Chemistry Gases MCQS

Taken from Inter part

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Gases MCQS Taken from Inter part

Q.1 The order of the rate of diffusion of gases NH_3 , SO_2 , Cl_2 and CO_2 is:

- (a) $\text{NH}_3 > \text{SO}_2 > \text{Cl}_2 > \text{CO}_2$
- (b) $\text{NH}_3 > \text{CO}_2 > \text{SO}_2 > \text{Cl}_2$
- (c) $\text{Cl}_2 > \text{SO}_2 > \text{CO}_2 > \text{NH}_3$
- (d) None of these

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Q.2 Pressure remaining constant at which temperature the volume of

gas will become twice of what it is at 0C.

(a) 546 oC (b) 200 oC

(c) 546 K (d) 273 K

Q.3 Equal masses of methane and oxygen are mixed in an empty container at 25oC. The fraction of the total pressure exerted by the

oxygen is:

(a) (b)

(c) (d)

Q.4 Which of the following of will have the same number of molecules at STP?

(a) 280 cm³ of CO₂ and 280 cm³ of N₂O

(b) 11.2 dm³ of O₂ and 32 g of O₂

(c) 44g of CO₂ and 11.2 dm³ of CO

(d) 28g of N₂ and 5.6 dm³ of oxygen

Q.5 Number of molecules in one dm³ of water is close to:

(a) $\times 10^{23}$ (b) $\times 10^{23}$

(c) $\times 10^{23}$ (d) $55.6 \times 6.02 \times 10^{23}$

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Q.6 If absolute temperature of a gas is doubled and the pressure is

reduced to one half the volume of gas

will.

(a) remain unchanged (b) increases four times

(c) reduce to (d) be double

Q.7 How should the conditions be changed to prevent the volume of

a given mass of gas from expanding when its mass is increased

(a) temperature is lowered and pressure is increased

(b) temperature is increased and the pressure is lowered

(c) temperature and pressure both are lowered

(d) temperature and pressure both are increased

Q.8 The molar volume of CO₂ is maximum at

(a) STP

(b) 127 °C and 1 atm

(c) 0 °C and 2 atm

(d) 273 °C and 2 atm

Q.9 Gases deviate from ideal behaviour at high pressure. Which of

the following is correct for non-ideality?

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(a) At high pressure, the gas molecules move in one direction only

(b) At high pressure, the collisions between, the gas molecules are

increased manifold

(c) At high pressure, the volume of gas becomes insignificant

(d) AT high pressure, the intermolecular attractions, become significant

Q.10 The deviation of a gas from ideal behaviour is maximum at,

(a) – 10 oC and 5.0 atm (b) – 10 oC and 2.0 atm

(c) 100 oC and 2.0 atm (d) 0 oC and 2.0 atm

Q.11 At high temperature isotherm moves away from both the axis

because of increase in,

(a) pressure (b) volume

(c) no. of moles (d) all above

Q.12 Values of Charle's law constant K depends upon.

(a) mass of gas (b) pressure gas

(c) no. of moles of gas (d) all above

Q.13 Equal volumes of H₂ and He are inserted in the same vessel. The

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pressure exerted by H₂ and He are in the ratio:

- (a) 1:1 (b) 2:1
- (c) 1:2 (d) all above

Q.14 Which of the following have same no. of molecules at STP

- (a) 1000 cm³ of N₂H₄ and O₂
- (b) 200 cm³ of CO₂ and N₂O
- (c) 50 cm³ each of CO and N₂
- (d) all above

Q.15 If absolute temperature is doubled and the pressure is increased 4

times. The volume is

- (a) half (b) double
- (c) four times (d) remains the changed

Q.16 Density of a gas is usually expressed in

- (a) kg m³ (b) kg dm³
- (c) g dm⁻³ (d) g cm⁻³

Q.17 Units of gas constant R in SI system is:

- (a) 0.0821 dm³ atm k⁻¹ mol⁻¹
- (b) 82.1 cm³ atm k⁻¹
- (c) 8.31 Nm k⁻¹ mol⁻¹
- (d) 1.987 cal k⁻¹ mol⁻¹

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Q.18 Concept of distribution of velocities among the gas molecules

was developed by

- (a) Claudius (b) Maxwell
- (c) Boltzman (d) Vanderwaal

Q.19 Absolute temperature of gas is proportional to

- (a) translational kinetic energy
- (b) rotational kinetic energy
- (c) vibrational kinetic energy
- (d) potential energy

Q.20 Deviation a gas from ideal behaviour is maximum at

- (a) low temperature, low pressure
- (b) low temperature, high pressure
- (c) high temperature, low pressure
- (d) high temperature high pressure

Q.21 Most ideal gas at room temperature is:

- (a) CO₂ (b) NH₃
- (c) SO₂ (d) N₂

Q.22 22.414 dm³ of various ideal gases at STP will have Avogadro's

number of molecules

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- (a) 6.02×10^{23} (b) 6.02×10^{24}
(c) 0.602×10^{23} (d) 6.02×10^{22}

Q.23 Gases are ideal at

- (a) low pressure and high temperature
(b) low temperature and high pressure
(c) high pressure and high temperature
(d) low pressure and low temperature

Q.24 The value of compressibility factor for an ideal gas is equal to:

- (a) 1 (b) 1.5
(c) 2 (d) 2.5

Q.25 An ideal gas obeys

- (a) Boyle's law (b) Charle's law
(c) Avogadro's law (d) all above

Q.26 A real gas obeying Vander Waal's equation will resemble ideal

gas if:

- (a) both 'a' and 'b' are large
(b) both 'a' and 'b' are small
(c) 'a' is small and 'b' is large
(d) 'a' is large and 'b' is small

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Q.27 Deep sea divers breath mixture of nitrogen and oxygen in a ratio

of:

(a) 96% N₂ and 4% O₂ (b) 4% and N₂ and 96% O₂

(c) 80% N₂ and 20% O₂ (d) 20% N₂ and 80% O₂

Q.28 One mole of any gas at STP occupies volume

(a) 2.24 dm³ (b) 22.4 dm³

(c) 44.4 dm³ (d) 48.4 dm³

Q.29 K.E. of gas molecules is equal to:

(a) (b) $m v^2$

(c) (d)

Q.30 All gases solidify before reaching at

(a) 373 oK (b) 273 oC

(c) - 473 oC (d) 0 oK

Q.31 Kinetic equation is equal to

(a) $P V = n R T$ (b) $P V = R T$

(c) $P V = m n c^2$ (d) $P V = m n c^2$

Q.32 Root mean square velocity is equal to

(a) (b)

(c) (d)

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Q.33 Kinetic energy associated with one molecule of a gas due to translational motion is given by

- (a) $E_k = \frac{1}{2} m v^2$ (b) $E_k = m v^2$
(c) $E_k = m n c^2$ (d) $E_k =$

Q.34 Density of gas is usually expressed as

- (a) kg m^{-3} (b) kg dm^{-3}
(c) g dm^{-3} (d) g cm^{-3}

Q.35 Weight of one dm^3 of O_2 at STP is

- (a) 1.4384 gm (b) 1.4394 gm
(c) 1.6384 gm (d) 1.3384 gm

Answers

Question 1 2 3 4 5

Answers b c a a a

Questions 6 7 8 9 10

Answers b a b d a

Questions 11 12 13 14 15

Answers b d a d a

Questions 16 17 18 19 20

Answers c c b a b

Questions 21 22 23 24 25

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Answers d a a a d

Questions 26 27 28 29 30

Answers b a b d d

Questions 31 32 33 34 35

Answers c a a c a

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