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Chemistry Mcqs from Gases, Liquids & Solids Solved for Lecturer

2020 PPSC Lecturer Preparation

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Chemistry Mcqs from Gases, Liquids & Solids for Lecturer Preparation

Q.1 The order of the rate of diffusion of gases NH_3 , SO_2 , Cl_2 and

CO_2 is:

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- (a) $\text{NH}_3 > \text{SO}_2 > \text{Cl}_2 > \text{CO}_2$
(b) $\text{NH}_3 > \text{CO}_2 > \text{SO}_2 > \text{Cl}_2$
(c) $\text{Cl}_2 > \text{SO}_2 > \text{CO}_2 > \text{NH}_3$
(d) None of these

Q.2 Pressure remaining constant at which temperature the volume of

gas will become twice of what it is at 0°C .

- (a) 546°C (b) 200°C
(c) 546 K (d) 273 K

Q.3 Equal masses of methane and oxygen are mixed in an empty

container at 25°C . The fraction of the total pressure exerted by the

oxygen is:

- (a)** (b)
(c) (d)

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Q.4 Which of the following of will have the same number of molecules at STP?

- (a) **280 cm³ of CO₂ and 280 cm³ of N₂O**
- (b) 11.2 dm³ of O₂ and 32 g of O₂
- (c) 44g of CO₂ and 11.2 dm³ of CO
- (d) 28g of N₂ and 5.6 dm³ of oxygen

Q.5 Number of molecules in one dm³ of water is close to:

- (a) **6.02×10^{23}**
- (b) 6.02×10^{23}
- (c) 6.02×10^{23}
- (d) $55.6 \times 6.02 \times 10^{23}$

Q.6 If absolute temperature of a gas is doubled and the pressure is

reduced to one half the volume of gas

will.

- (a) remain unchanged
- (b) **increases four times**
- (c) reduce to
- (d) be double

Q.7 How should the conditions be changed to prevent the volume of

a given mass of gas from expanding when its mass is increased

- (a) temperature is lowered and pressure is increased**
- (b) temperature is increased and the pressure is lowered
- (c) temperature and pressure both are lowered
- (d) temperature and pressure both are increased

Q.8 The molar volume of CO₂ is maximum at

- (a) STP
- (b) 127 oC and 1 atm**
- (c) 0 oC and 2 atm
- (d) 273 oC and 2 atm

Q.9 Gases deviate from ideal behaviour at high pressure.

Which of

the following is correct for non-ideality?

- (a) At high pressure, the gas molecules move in one direction only
- (b) At high pressure, the collisions between, the gas molecules are

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increased manifold

(c) At high pressure, the volume of gas becomes insignificant

(d) AT high pressure, the intermolecular attractions, become significant

Q.10 The deviation of a gas from ideal behaviour is maximum at,

(a) - 10 oC and 5.0 atm (b) - 10 oC and 2.0 atm

(c) 100 oC and 2.0 atm (d) 0 oC and 2.0 atm

Q.11 At high temperature isotherm moves away from both the axis

because of increase in,

(a) pressure (b) **volume**

(c) no. of moles (d) all above

Q.12 Values of Charle's law constant K depends upon.

(a) mass of gas (b) pressure gas

(c) no. of moles of gas (d) **all above**

Q.13 Equal volumes of H₂ and He are inserted in the same vessel. The

pressure exerted by H₂ and He are in the ratio:

- (a) **1:1** (b) 2:1
(c) 1:2 (d) all above

Q.14 Which of the following have same no. of molecules at STP

- (a) 1000 cm³ of N₂H₄ and O₂
(b) 200 cm³ of CO₂ and N₂O
(c) 50 cm³ each of CO and N₂
(d) **all above**

Q.15 If absolute temperature is doubled and the pressure is increased 4

times. The volume is

- (a) **half** (b) double
(c) four times (d) remains the changed

Q.16 Density of a gas is usually expressed in

- (a) kg m³ (b) kg dm³

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(c) **g dm⁻³** (d) g cm⁻³

Q.17 Units of gas constant R in SI system is:

(a) 0.0821 dm³ atm k⁻¹ mol⁻¹

(b) 82.1 cm³ atm k⁻¹

(c) **8.31 Nm k⁻¹ mol⁻¹**

(d) 1.987 cal k⁻¹ mol⁻¹

Q.18 Concept of distribution of velocities among the gas molecules

was developed by

(a) Claudius

(b) **Maxwell**

(c) Boltzman

(d) Vanderwaal

Q.19 Absolute temperature of gas is proportional to

(a) **translational kinetic energy**

(b) rotational kinetic energy

(c) vibrational kinetic energy

(d) potential energy

Q.20 Deviation a gas from ideal behaviour is maximum at

(a) low temperature, low pressure

(b) **low temperature, high pressure**

(c) high temperature, low pressure

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(d) high temperature high pressure

Q.21 Most ideal gas at room temperature is:

- (a) CO₂ (b) NH₃
(c) SO₂ (d) **N₂**

Q.22 22.414 dm³ of various ideal gases at STP will have Avogadro's

number of molecules

- (a) **6.02 x 10²³** (b) 6.02 x 10²⁴
(c) 0.602 x 10²³ (d) 6.02 x 10²²

Q.23 Gases are ideal at

- (a) **low pressure and high temperature**
(b) low temperature and high pressure
(c) high pressure and high temperature
(d) low pressure and low temperature

Q.24 The value of compressibility factor for an ideal gas is equal to:

- (a) **1** (b) 1.5
(c) 2 (d) 2.5

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Q.25 An ideal gas obeys

- (a) Boyle's law (b) Charle's law
(c) Avogadro's law (d) **all above**

Q.26 A real gas obeying Vander Waal's equation will resemble ideal

gas if:

- (a) both 'a' and 'b' are large
(b) **both 'a' and 'b' are small**
(c) 'a' is small and 'b' is large
(d) 'a' is large and 'b' is small

Q.27 Deep sea divers breath mixture of nitrogen and oxygen in a ratio

of:

- (a) **96% N₂ and 4% O₂** (b) 4% and N₂ and 96% O₂
(c) 80% N₂ and 20% O₂ (d) 20% N₂ and 80% O₂

Q.28 One mole of any gas at STP occupies volume

- (a) 2.24 dm³ (b) **22.4 dm³**

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(c) 44.4 dm³ (d) 48.4 dm³

Q.29 K.E. of gas molecules is equal to:

(a) (b) $m v^2$

(c) (d)

Q.30 All gases solidify before reaching at

(a) 373 oK (b) 273 oC

(c) - 473 oC (d) 0 oK

Q.31 Kinetic equation is equal to

(a) $P V = n R T$ (b) $P V = R T$

(c) **$PV = m n c^2$** (d) $PV = m n c^2$

Q.32 Root mean square velocity is equal to

(a) (b)

(c) (d)

Q.33 Kinetic energy associated with one molecule of a gas due to

translational motion is given by

(a) **$E_k =$** (b) $E_k = m v^2$

(c) $E_k = m n c^2$ (d) $E_k =$

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Q.34 Density of gas is usually expressed as

- (a) kg m⁻³ (b) kg dm⁻³
(c) **g dm⁻³** (d) g cm⁻³

Q.35 Weight of one dm³ of O₂ at STP is

- (a) **1.4384 gm** (b) 1.4394 gm
(c) 1.6384 gm (d) 1.3384 gm

Q.36 Ionic solids are characterized by

- (a) low melting points
(b) good conductivity in solid state
(c) high vapour pressure
(d) **solubility in polar solvents**

Q.37 Amorphous solids.

- (a) have sharp melting points
(b) undergo clean cleavage when cut with knife
(c) have perfect arrangements of atoms
(d) **can presses small regions of orderly arrangements of atoms**

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Q.38 The force of attraction between the atoms of helium is

- (a) hydrogen bonding
- (b) coordinate covalent bond
- (c) covalent bond
- (d) London dispersion force**

Q.39 Which of the following is a pseudo-solid

- (a) CaF_2
- (b) Glass**
- (c) NaCl
- (d) All

Q.40 Diamond is a bad conductor because

- (a) It has a tight structure
- (b) It has a high density
- (c) There is no free electron present in the crystal of diamond to**

conduct electricity

- (d) None of the above

Q.41 The weakest intermolecular force is

- (a) dipole-dipole force
- (b) electrostatic force between ions
- (c) ion-dipole force

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(d) dipole-induced dipole force

Q.42 In liquids intermolecular forces are

(a) very weak (b) very strong

(c) reasonably strong (d) ion-dipole force

Q.43 Values of heat of vaporization for liquids, with strong dipole-

dipole forces will be

(a) very high (b) very low

(c) reasonably high (d) negligible

Q.44 Instantaneous dipole-induced dipole force is also called

(a) dipole force **(b) London dispersion**

(c) hydrogen bonding (d) none of the above

Q.45 Down the group polarizability generally

(a) increases (b) decreases

(c) remains constant (d) do not follow a regular trend

Q.46 Trend of boiling points of halogens from fluorine to iodine is

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that it.

- (a) decreases (b) **increases**
(c) remains constant (d) negligible

Q.47 Molecules of hydro carbons with large chain lengths experience

- (a) repulsive forces (b) **strong attractive force**
(c) weaker attractive forces
(d) no attractive force

Q.48 Hydrocarbons which generally have high molecular masses exist

is.

- (a) **solid form** (b) liquid form
(c) vapour form (d) gaseous form

Q.49 Exceptionally low acidic strength of HF is due to

- (a) strong polar bond (b) small size of fluorine
(c) **strong hydrogen bonding** (d) Vander Waal's forces

Q.50 Long chain of amino acids are coiled about one another into spiral by.

- (a) covalent bond (b) ionic bond
(c) hydrogen bond (d) Vander Waal's forces

Q.51 Evaporation of water is possible at

- (a) 100°C** (b) 0°C
(c) at all temperatures (d) above 100°C

Q.52 Boiling point is low for liquid with

- (a) high vapour pressure at given temperature
(b) low vapour pressure at a given temperature
(c) very high vapour pressure
(d) very low vapour pressure

Q.53 At equilibrium rate of evaporation and rate of condensation

- (a) become very high (b) become very low
(c) can never be equal **(d) become equal**

Q.54 In an open system vapour pressure of water at 100°C at sea level

is

- (a) 700 mm of Hg (b) **760 mm of Hg**
(c) 670 mm of Hg (d) 1000 mm of Hg

Q.55 Molar heat of vaporization of water is

- (a) 140.6 kJ/mol (b) 14.06 kJ/mol
(c) 18 kJ/mol (d) **40.6 kJ/mol**

Q.56 When external pressure is 23.7 torr boiling point of water is

- (a) 100°C (b) 200°C
(c) 98°C (d) **25°C**

Q.57 Distillation under very reduced pressure is called

- (a) fractional distillation (b) distillation
(c) **vacuum destructive distillation**
(d) destructive distillation

Q.58 Water may boil at 120°C when external pressure is

- (a) 760 torr (b) 100 torr
(c) 1489 torr (d) 700 torr

Q.59 Amount of heat absorbed when one mole of solid melts into

liquid form at its melting point is called

- (a) molar heat of sublimation
(b) heat of vaporization
(c) latent heat of fusion
(d) molar heat of fusion

Q.60 Ethanol is much more soluble in water than ethyl ethanoate

which one of the following statements correctly account for this

- (a) ethanol is polar molecule but ethyl ethanoate is non-polar
(b) ethanol is non polar molecule but ethyl ethanoate is polar
(c) a hydrogen bond is formed between H-atom of the OH group in

ethanol and O-atom of water molecule

- (d) a hydrogen bond is formed between the H-atom of the OH

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group in ethanol and hydrogen of the water molecule

Q.61 The boiling point of a liquid will be

- (a) lower at high altitude**
- (b) higher at high altitude
- (c) same at sea level and high altitudes
- (d) equal to atmospheric pressure

Q.62 The process in which liquids can be made to boil at low temperature is called

- (a) vacuum distillation**
- (b) destructive distillation
- (c) distillation
- (d) vacuum destructive distillation

Q.63 Why is the boiling point of methane greater than that of neon

- (a) a molecule of methane has a greater mass
- (b) a molecule of methane has more electrons than a molecule of neon

(c) the molecules of methane have stronger intermolecular forces

than those of neon

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(d) the molecule of methane is polar but that of neon is not

Q.64 The amount of heat required to vaporize one mole of a liquid at

its boiling point is called

(a) molar heat of vaporization

(b) molar heat of fusion

(c) latent heat of fusion

(d) molar heat of sublimation

Q.65 Which of the elements in its crystalline form will have the lowest

enthalpy change of vaporizations

(a) chlorine **(b) argon**

(c) phosphorous (d) silicon

Q.66 Crystals show variation in physical properties depending upon

the direction. The property is called

(a) isomorphism (b) polymorphism

(c) anisotropy (d) isotropy

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Q.67 Certain melt to a turbid liquid phase with properties of liquids as

well as some degree of order like solid. Such turbid liquids are called

- (a) anorphous solid (b) vitreous solid
(c) crystalline solid (d) **liquid crystal**

Q.68 Isomorphous crystals show

- (a) same chemical properties
(b) same physical properties
(c) **same crystalline form**
(d) same melting point

Q.69 Existence of an element in more than form is known as

- (a) **allotropy** (b) isomorphism
(c) isotropy (d) none of these

Q.70 Crystalline forms of the same, substance can coexist in equilibrium with each other at its

- (a) melting point (b) **transition temperature**
(c) boiling point (d) none of these

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Q.71 Crystal lattice of substance can be categorised into

- (a) five types (b) **seven types**
(c) six types (d) none of these

Q.72 Covalent solids are composed of

- (a) ions (b) different molecules
(c) **neutral atoms** (d) any of the above

Q.73 Carbon atoms of diamond are

- (a) sp hybridized (b) sp² hybridized
(c) **sp³ hybridized** (d) unhybridized

Q.74 Molecular crystals are generally

- (a) hard (b) **soft**
(c) unstable (d) stable

Q.75 Ionic crystals are

- (a) **hard** (b) soft
(c) brittle (d) amorphous

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