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PPSC CHEMISTRY PAPER

LECTURER 2015 MCQS

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1. Enzymes belong to which class of compound-----Proteins----- right answer)
2. The helical structure of protein is stabilized by -----Hydrogen-----bond. right answer)
3. Enzymes have...Polypeptide.....unit
A) Polypeptide right answer)
B) HYDROCARBON
4. Function of enzymes in biological system is.....

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Catalyze biochemical action right answer)

5. Night blindness is produce due to deficiency of ...

.....VITAMIN A right answer)

6. Zwitter ion exist in which of the following

a) Alanine right answer)

b) Glycine HCL

c) Both a and b

d) None

7. IR determine -----bond

a) kind of bond nature of bond right answer)

c) hydrocarbon

8. Marsh gas is

...CH₄ right answer)

9. Solvent used in IR Spectroscopy is.....

a) CCl₄ right answer)

B) CH₃OH

C) WATER

D) CH₃CH₂OH

10. 1700-----1760 PEAK In IR Spectroscopy is.....

...Carbonyl group right answer)

11. NMR use -----source of radiation

a) UV

B) Radio waves right answer)

c) micro waves

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d) visible rays

12. Existence of nucleus is confirmed by.....

Millikian method or

Alpha scattering by thin metal foil right answer)

13. What is false about electron (Not conform but it should be d)

a) Electron is a particle

b) Wave is associated with electron

c) When electron jump from orbital its energy c..

d) oriented by magnetic field

14. Which one is occur in whole number always

a) Atomic number right answer)

b) Molecular weight

c) Equivalent weight

15. A metal has an electronic configuration is $M+2$ has 2,8,14 and its atomic weight is 56. the number of neutron is

.....

30) right answer)

16. Least soluble in Water Is

a. Ethano

b. I b. Phenol

c. c. Carboxylic acid

d. d. benzene right answer)

17. The EAN Value of $Ni(CO)_4$ is

a) 36 right answer)

b) 54

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18. The ratio of C:Si in their neutron ratio with atomic weight having 12 and 28 respectively

(3:7) right answer)

19. Phenolphthalein is colourless in acidic medium due to ___
___ Benzenoid form right answer)

20. The weakest base is

- a) H-
- b) Cl- right answer)
- c) HCO₃⁻
- d) OH-

21. Which compound decolorize KMnO₄ solution but not give test with AgNO₃ solution a)

- Ethylene right answer)
- b) Acetylene
- c) Ethane

22. Which one of the following does not give Lassaignes test for nitrogen

- A. Phenylhydrazine
- B. Glycine
- C. Urea
- D. Azobenzene right answer)

23. which one the following compound does not explode:

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- a) trinitro toluene
- b) 2-amine toluene (I think this is not conform)

24. Strongest oxidizing agent is

supported by-----

- a. Low dissociation energy
- b. high electron affinity right answer)
- c. low ionization energy

25. the order of oxidizing agent is

$I < Br < Cl < F$

26. inversion of cane sugar into glucose and fructose is ?(not conform)

- a) slow process fast process
- b) spontaneous process instantaneous process

27. rate of chemical reaction is-----when it proceed

- a. decrease right answer)
- b. increase

28. when reactants are added then rate of reaction is (not conform) decrease

increase right answer)

29. which one of the following metal has highest electrical conductivity

- A. gold
- B. silver right answer)

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- C. aluminium
- D. copper

30. preparation of ice cream is

- a) colloids
- b) excessive cooling (right answer)
- c) coagulation
- d) peptization

31. colloids are example of -----equilibria

- a. homogenous equilibria
- A. heterogenous equilibria (right answer)
- b. suspension
- c. none of these

32. pure water is nonconductor of electricity

- A. water is neutral
- B. water is almost unionized (right answer)
- B. water is completely ionized

33. which one of the following aqueous solution conductor of electricity

- C. Glucose
- D. HCl (right answer)

34. Which is correct statement (not confirm my answer is c)

- A. Catalyst increase the rate of forward reaction
- B. Catalyst increase the rate of backward reaction
- C. Catalyst influence on the rate of forward reaction and backward reaction equally (right answer)
- D. Catalyst decrease the rate of forward reaction

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35. Pick a right statement:

Order of reaction is in fraction (right answer)

36. The unit for formation of ammonia from



Mol⁻².lit²(right answer)

37. The unit of active mass is----- A. Gram atom per litre (right answer)

B. Gram molecule per litre

C. Gram litre

38. colloidal particles are

a) negative

b) positive

c) neutral (my answer is neutral not confirm)

d) may be positive or negative

39. who classify element first

a) lother meyer

b) doberienner

c) newlands

d) mendleevs

40. molecularity of reaction is always inwhole number (right answer)

41. the chemical which destroy ozone layer is CFCI₃ (right answer)

42. Physisorption occur at -----temperature low (right answer)

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43. if dispersion medium is gas its example is

Aerosol(right answer)

44. pressure vs temperature equation is called (not conform answer)

- A. Langmuir equation
- B. Fredlucck equation
- C. Bet equation
- D. None of these

45. Resonating structure caannot be seprated (right answer)

46. Nickel is use in the hydrogenation of vegetable ghee (right answer)

47. Colloids shows Brownian movement(right answer)

48. Methyl group is Ortho and p-directiong bcz.....

(Not conform my answer is all ob above)

- a) E donating effect of methyl group inductively
- b) E donating effect of methyl group resonatingly
- c) E donating effect of methyl group hyperconjugation (right answer)
- d) All of above

49. Chloro group is ortho and para but electron withdrawing (right answer)

50. To avoid maximum yield For a Reversible reaction to occur

- A. Glass vessel
- B. Closed vessel(right answer)
- C. Open vessel

51. In benzene, there are -----type of carbon atom

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- A. One type (right answer)
- B. Two type
- C. Three type
- D. Four type

52. In benzene, the number of pi electron is

- A. 3 e
- B. 4e
- C. 6e (right answer)
- D. 4e

53. In benzene , all carbon atoms are-----hybridization

- A. Sp³
- B. Sp² (right answer)
- C. Sp
- D. None

54. According to lechatlier principle in a reversible reaction between solid and liquid, the amount of heat

added

(Not conform answer)

- A. Decrease the concentration of solid
- B. increase the concentration of solid
- C. Decrease the concentration of liquid
- D. Decrease the concentration of liquid

55. The fastest reaction occur in which of the following:

- A. Rusting of iron
- B. Burning of coal
- C. The AgCl formation by the reaction of AgNo₃ and Nacl (right answer)

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56. The amount of energy required to separate nucleons is called

- A. Ionization energy
- B. Binding energy (right answer)

57. The enthalpy of heat of formation of sodium ion from its atom is known as

- A. Enthalpy of atomization
- B. First ionization energy (not conform)
- C. Enthalpy of combustion

58. The brown vapour is formed of which element when MnO_2 and conc. H_2SO_4 is added in them A. NO_2

(right answer)

- B. Cl_2
- C. Br_2
- D. I_2

59. Stability of nucleus is due to

- A. Proton and neutron (right answer)
- B. Proton and electron
- C. Neutron and electron

60. Difference between particle mass and real mass is called.....

Mass defect (right answer)

61. A gas has a volume of 2 litre at S.T.P at constant pressure, the new volume is 4 litre at which

temperature: A. $50^\circ C$

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B. 100c

62. The difference between the crystalloids and colloids

- A. Particle size (right answer)
- B. Colloidal size
- C. Diffusion
- D. True solution

63. The gas which shows deviation from ideal behavior is:

- A. Helium
- B. Hydrogen
- C. Ammonia (right answer)
- D. Trichloro methane

64. Some substance shows scattering of light but they can pass through ordinary filter paper is.....

- A. Colloids (right answer)
- B. Suspension
- C. True solution
- D. None

65. In the lime kiln , $\text{CaCO}_3 \dots \dots \dots \rightarrow \text{CaO} + \text{CO}_2$ proceed to complete because.....

- A. CO_2 ESCAPPE
- B. CaO remove
- C. Low Temperature
- D. High Pressure

66. Which one of the following equation is not affected by change in pressure

- A. $\frac{1}{2}\text{N}_2 + \frac{1}{2}\text{O}_2 \dots \dots \dots \rightarrow \text{NO}$ (right answer)

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67. A Reaction approaches go farthest completions

A. $K=10^2$

B. $K=10^1$

C. $K=10$

D. $K=10^{-2}$

68. Isothermal thermal system is one in which

$\Delta T=0$ (right answer)

$\Delta E=W$

$V=0$

$E=0$

69. Temperature of the system decreases in adiabatic expansion (right answer) adiabatic compression

70. System in which no change in temperature is Isochoric isobaric

isothermic (right answer)

71. Cyclic structure of benzene is given by

Kekule(right answer)

Doberiner

72. Which is least soluble in water

A. Ethanol

B. Benzene (right answer)

C. Benzoic acid

D. Phenol

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73. Heterogenous catalysis is based on

- a. Adsorption (right answer)
- A. Absorption
- B. Sorption
- C. Desorption

74. Entropy of the universe is

- Always increasing (right answer)
- Decreasing
- Constant

75. Which indicator is not used in conical flask

- A. Self indicator
- B. External indicator (right answer)
- C. Internal indicator
- D. Mixed indicator

76. Toluene is prepared from benzene

- A. Friedel craft reaction (right answer)
- B. Perkin reaction
- C. Wurtz reaction

77. Catalyst used in Friedel craft alkylation is

- A. $AlCl_3$ (right answer)
- B. $BFeCl_3$

78. Order of reaction is.....

- Can be determined from equation
- Can be determined from experiment (right answer)

79. Oxidation number of P in KH_2PO_2 is

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+1 (right answer)

+3

+5

+2

80. Molten KCl conduct electricity because of

Free ions (right answer)

Free electron

81. which individual units loose identity

a. Complex salts right answer)

b. Simple salts

c. Compound salts

d. Double Salt

82. When a multidentate ligand is surrounded it is known as

Coordination sphere

Coordination complex

83. Group no. represents Same electronic configuration

same number of electron in valance shell(right answer) Chelates (right answer)

84. Entropy decreases in melting of ice rusting of iron

85. In this reaction $6\text{NaOH} + 3\text{Cl}_2 \dots\dots\dots 5\text{NaCl} + \text{NaClO}_3 + 3\text{H}_2\text{O}$

A. Chlorine is reduced

B. Chlorine is oxidized

C. Disproportionation reaction

86. Aqueous solution is electrical conductor MgCl_2

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87. According to Lewis concept, bases are electron pair donor

88. According to Lowry-Bronsted concept, bases are proton acceptor

89. Which one of the following aqueous solutions of oxide and chloride give

same test with litmus solution. A. Phosphorous

B. Sodium

C. Magnesium

90. The element of group 1B is known as coinage metal.

91. Oxidation of which gives positive DNP but not Fehling's test

A. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$

B. $\text{CH}_3\text{CH}(\text{OH})\text{CH}_2\text{CH}_3$

C. $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{OH}$

D. $(\text{CH}_3)_3\text{COH}$

92. Product of hydrolysis of $\text{CH}_3\text{CO}_2\text{C}_3\text{H}_7$ with NaOH gives

A. $\text{CH}_3\text{COONa} + \text{C}_3\text{H}_7\text{OH}$

93. Fluorescein, rhodamine, are examples of----- indicators.

A. Acid base indicator

B. Adsorption indicator

C. Absorption indicator

94. Evaporation causes a change in entropy.....

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