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# PPSC, FPSC (Lecturer- Chemistry) Solved Mcqs Preparation Notes-Part-3 On doc4shares.com

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**101. QUESTION:** The atomic mass of carbon is

\_\_\_\_\_.

- A. 11.0107
- B. 12.0107
- C. 13.0107
- D. 14.0107

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**CORRECT** ANSWER IS: 12.0107

**102. QUESTION: The electro-negativity of carbon is**

\_\_\_\_\_.

- A. 3.44
- B. 3.04
- C. 2.55
- D. 2.65

**CORRECT** ANSWER IS: 2.55

**103. QUESTION: Carbon is discovered in** \_\_\_\_\_.

- A. 3750 BC
- B. 3650 BC
- C. 3550 BC
- D. 1772

**CORRECT** ANSWER IS: 3750 BC

**104. QUESTION: The electronic configuration of carbon is**

\_\_\_\_\_.

- A.  $2s^2$
- B.  $2s^2 2p^2$
- C.  $2s^2 2p^3$
- D.  $2s^2 2p^4$

**CORRECT** ANSWER IS:  $2s^2 2p^2$

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**105. QUESTION: The state of Carbon is \_\_\_\_\_.**

- A. Solid
- B. Liquid
- C. Gass
- D. Plasma

**CORRECT** ANSWER IS: Solid

**106. QUESTION: The Melting point of carbon is \_\_\_\_\_.**

- A. 3550o C or 3823 K
- B. 3823o C or 3550 K
- C. 3000o C or 3000 K
- D. None of the above

**CORRECT** ANSWER IS: 3550o C or 3823 K

**107. QUESTION: the boiling point of carbon is \_\_\_\_\_.**

- A. 3825oC or 4098 K
- B. 3000oC or 4000 K
- C. 3800oC, 4500 K
- D. None of the above

**CORRECT** ANSWER IS: 3825oC, 4098 K

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**108. QUESTION: Helium is a \_\_\_\_\_.**

- A. Nobel gass
- B. Metal
- C. Non metal
- D. both a and c

**CORRECT** ANSWER IS: both a and c

**109. QUESTION: The symbol of helium is \_\_\_\_\_.**

- A. Hs
- B. HI
- C. He
- D. Hf

**CORRECT** ANSWER IS: He

**110. QUESTION: the atomic weight of helium is \_\_\_\_\_.**

- A. 3.003
- B. 4.003
- C. 5.003
- D. 6.003

**CORRECT** ANSWER IS: 4.003

**111. QUESTION: The state of helium is \_\_\_\_\_.**

- A. Gas

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- B. Liquid
- C. Plasma
- D. Solid

**CORRECT** ANSWER IS: Gas

**112. QUESTION: The boiling point of helium is**

\_\_\_\_\_.

- A. 268.9oC or 4.2 K
- B. -268.9oC or 5 K
- C. -268.9oC or 4.2 K
- D. 268.9oC or 5 K

**CORRECT** ANSWER IS: -268.9oC or 4.2 K

**113. QUESTION: The Melting point of helium is**

\_\_\_\_\_.

- A. 272.2oC, 95 K
- B. -272.2oC, 0.95 K
- C. -272.2oC, 95 K
- D. None of the above

**CORRECT** ANSWER IS: -272.2oC, 0.95 K

**114. QUESTION: The electronic configuration of helium is**

\_\_\_\_\_.

- A. 1s<sup>2</sup>
- B. 1s<sup>2</sup>2p<sup>2</sup>

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C.  $1s^3$

D.  $1s^2 2s^2 2p^3$

**CORRECT** ANSWER IS:  $1s^2$

**115. QUESTION: Number of electrons in helium are**

\_\_\_\_\_.

A. 2

B. 3

C. 4

D. 5

**CORRECT** ANSWER IS: 2

**116. QUESTION: Number of protons in helium are**

\_\_\_\_\_.

A. 2

B. 3

C. 4

D. 5

**CORRECT** ANSWER IS: 2

**117. QUESTION: Aluminum is a silver colored**

\_\_\_\_\_.

A. Metal

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- B. Non-metal
- C. Gas
- D. None of the above

**CORRECT** ANSWER IS: Metal

**118. QUESTION: Atomic weight of Aluminum is**

\_\_\_\_\_.

- A. 27
- B. 26.98
- C. 26.89
- D. 26

**CORRECT** ANSWER IS: 26.98

**119. QUESTION: There are \_\_\_\_\_ numbers of electrons in aluminum.**

- A. 13
- B. 14
- C. 15
- D. 16

**CORRECT** ANSWER IS: 13

**120. QUESTION: Electronic configuration of Aluminum is**

\_\_\_\_\_.

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- A.  $2s^2 2p^6 3s^2 3p^1$
- B.  $1s^2 2s^2 2p^6 3s^2 3p^1$
- C.  $1s^2 2s^2 2p^6 3s^2$
- D.  $1s^2 2s^2 2p^6 3p^1$

**CORRECT** ANSWER IS:  $1s^2 2s^2 2p^6 3s^2 3p^1$

**121. QUESTION: Melting point of aluminum is**

\_\_\_\_\_.

- A. 660.32 oC, 933.57 K
- B. 560.32 oC, 933.57 K
- C. 460.32 oC, 933.57 K
- D. 360.32 oC, 933.57 K

**CORRECT** ANSWER IS: 660.32 oC, 933.57 K

**122. QUESTION: The boiling point of aluminum is**

\_\_\_\_\_.

- A. 2000 oC, 2740.00 K
- B. 2000 oC, 5000 K
- C. 2000 oC, 1000 K
- D. 2466.85oC, 2740.00 K

**CORRECT** ANSWER IS: 2466.85oC, 2740.00 K

**123. QUESTION: Lithium is a \_\_\_\_\_.**

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- A. Halogens
- B. Nobel gas
- C. Alkali metal
- D. Alkaline earth metal

**CORRECT** ANSWER IS: Alkali metal

**124. QUESTION: State of lithium is \_\_\_\_\_.**

- A. Solid
- B. Gas
- C. Plasma
- D. Liquid

**CORRECT** ANSWER IS: Solid

**125. QUESTION: \_\_\_\_\_ is a pseudo solid.**

- A. NaCl
- B. Glass
- C. Both a and b
- D. None of the above

**CORRECT** ANSWER IS: Glass

**126. QUESTION: In structure of NaCl, each ion is surrounded by \_\_\_\_\_ number of chloride ions.**

- A. 4

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B. 5

C. 6

D. 7

**CORRECT** ANSWER IS: 6

**127. QUESTION: The \_\_\_\_\_ of main group metal oxides increases on descending a group.**

A. Basicity

B. Amphotericity

C. Acidity

D. Liquidity

**CORRECT** ANSWER IS: Basicity

**128. QUESTION: Which ion has high affinity for proton and cannot exist alone in an aqueous solution?**

A. H +

B. O -2

C. H 2+

D. All of the above

**CORRECT** ANSWER IS: O-2

**129. QUESTION: The \_\_\_\_\_ of main group metal oxides increases on descending a group.**

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- A. Transition
- B. Oxidation
- C. Inert
- D. None of the above

**CORRECT** ANSWER IS: Oxidation

**130. QUESTION: Semi metals have properties of**

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- A. Metals
- B. belongs to III A, IV A V A
- C. Non-Metals
- D. All of the above

**CORRECT** ANSWER IS: All of the Above

**131. QUESTION: Hydrogen is recognized as distinct substance by \_\_\_\_\_ in 1776.**

- A. Joseph Priestley
- B. Carl Wilhelm Scheele
- C. Henry Cavendish
- D. None of the above

**CORRECT** ANSWER IS: Henry Cavendish

**132. QUESTION: Atomic Number of Hydrogen is**

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- A. 1
- B. 2
- C. 3
- D. 4

**CORRECT** ANSWER IS: 1

**133. QUESTION: Atomic symbol of Hydrogen is**

\_\_\_\_\_.

- A. H
- B. D
- C. R
- D. G

**CORRECT** ANSWER IS: H

**134. QUESTION: Atomic Weight of Hydrogen is**

\_\_\_\_\_.

- A. 1.008
- B. 1.009
- C. 1.010
- D. 1.011

**CORRECT** ANSWER IS: 1.008

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**135. QUESTION: Electronic Configuration of Hydrogen is**

\_\_\_\_\_.

- A.  $2s^1$
- B.  $1s^1$
- C.  $2s^2$
- D.  $3s^1$

**CORRECT** ANSWER IS:  $1s^1$

**136. QUESTION: Atomic Radius of hydrogen is**

\_\_\_\_\_.

- A. 117 pm
- B. 118 pm
- C. 119 pm
- D. 120 pm

**CORRECT** ANSWER IS: 120 pm

**137. QUESTION: Melting point of hydrogen is**

\_\_\_\_\_.

- A. 259.16 oC
- B. -252.879 oC
- C. -259.16 oC
- D. 252.8976 oC

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**CORRECT** ANSWER IS: -259.16 oC

**138. QUESTION: Boiling point of hydrogen is**

\_\_\_\_\_.

- A. 259.16 oC
- B. -252.879 oC
- C. -259.16 oC
- D. 252.8976 oC

**CORRECT** ANSWER IS: -252.879 oC

**139. QUESTION: Oxidation states of hydrogen is**

\_\_\_\_\_.

- A. -1, +1 an amphoteric oxide
- B. +1 Strongly basic Oxide
- C. +2, +1 an amphoteric oxide
- D. 1

**CORRECT** ANSWER IS: -1, +1 an amphoteric oxide

**140. QUESTION: Sodium is first discovered by**

\_\_\_\_\_.

- A. Joseph Priestley
- B. Carl Wilhelm Scheele

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C. Henry Cavendish

D. Humphry Davy

**CORRECT** ANSWER IS: Humphry Davy

**141. QUESTION: Sodium was first isolated by Humphry Davy in \_\_\_\_\_.**

A. 1806

B. 1807

C. 1808

D. 1809

**CORRECT** ANSWER IS: 1807

**142. QUESTION: Atomic number of sodium is**

\_\_\_\_\_.

A. 10

B. 11

C. 12

D. 13

**CORRECT** ANSWER IS: 11

**143. QUESTION: Atomic symbol of sodium is**

\_\_\_\_\_.

A. Nacl

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B. Na

C. S

D. H

**CORRECT** ANSWER IS: Na

**144. QUESTION: Atomic weight of sodium is**

\_\_\_\_\_.

A. 21.99

B. 22.99

C. 23.99

D. 24.99

**CORRECT** ANSWER IS: 22.99

**145. QUESTION: Electronic Configuration of sodium is**

\_\_\_\_\_.

A. 1s1

B. 2s1

C. 3s1

D. 3s1

**CORRECT** ANSWER IS: 3s1

**146. QUESTION: Atomic radius of sodium is**

\_\_\_\_\_.

A. 225 pm

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- B. 226 pm
- C. 227 pm
- D. 228 pm

**CORRECT** ANSWER IS: 227 pm

**147. QUESTION: Melting point of sodium is**

\_\_\_\_\_.

- A. 97.8 oC
- B. 96.8 oC
- C. 95.8 oC
- D. 94.8 oC

**CORRECT** ANSWER IS: 97.8 oC

**148. QUESTION: boiling point of Sodium is**

\_\_\_\_\_.

- A. 783 oC
- B. 883 oC
- C. 983 oC
- D. 1083 oC

**CORRECT** ANSWER IS: 883 oC

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**149. QUESTION: Oxidation states of sodium is**

\_\_\_\_\_.

- A. -1, +1 an amphoteric oxide
- B. +1 Strongly basic Oxide
- C. +2, +1 an amphoteric oxide
- D. 1

**CORRECT** ANSWER IS: 1

**150. QUESTION: Potassium is discovered by**

\_\_\_\_\_.

- A. Joseph Priestley
- B. Carl Wilhelm Scheele
- C. Henry Cavendish
- D. Humphry Davy

**CORRECT** ANSWER IS: Humphry Davy

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